

Crystal nucleation in silicate glasses: the temperature and size dependence of crystal/liquid surface energy

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Abstract

The most basic assumption of the classical nucleation theory (CNT) is to treat nucleus/liquid surface energy, σ , as a macroscopic property having a value equal to that of a planar interface, σ_∞ . Therefore, when the CNT is used to analyze experimental data, the size dependence of surface energy is often neglected. To date, there has been no reliable method to measure the surface energy of the nucleus/liquid interface except by fitting nucleation rate data to the theory. In this case, one obtains the surface energy of critical size nuclei as a function of temperature. However, the fitted $\sigma(T)$ dependence arises from two different factors: the temperature dependence of σ for a planar interface and its size dependence. This paper focuses on the temperature dependence of the macroscopic value of surface energy, decoupling it from the size dependent part. Tolman's equation was used to eliminate the size dependence of surface energy from published nucleation data for two stoichiometric silicate glasses ($\text{Li}_2\text{O} \cdot 2\text{SiO}_2$ and $\text{Na}_2\text{O} \cdot 2\text{CaO} \cdot 3\text{SiO}_2$). It is shown that the Tolman parameter may be chosen so that surface tension decreases with temperature; $d\sigma_\infty/dT < 0$. The value of $d\sigma_\infty/dT$ obtained in this way is close to theoretical predictions. © 2000 Elsevier Science B.V. All rights reserved.

1. Introduction

According to the classical nucleation theory (CNT) the nucleus/liquid interfacial energy, σ , governs the crystal nucleation rate to a large extent [1]. The most serious assumption of the CNT, however, is that the thermodynamic properties of nanometric aggregates of the newly nucleated phase are the same as those of the corresponding bulk phase. Thus, nucleus/liquid surface energy is

treated as a macroscopic property with a value equal to the respective value for a planar interface, σ_∞ . This assumption is known as the capillarity approximation.

Due to the lack of a reliable (direct) method to measure the surface energy of the nucleus/liquid interface, independently of nucleation experiments, one is forced to fit experimental data to theory. Hence, σ_∞ cannot be compared directly with experimental values.

When the CNT is used to treat experimental data, not only the size dependence of surface energy but also its temperature dependence is often neglected. The latter approximation allows for the calculation of σ_∞ and the pre-exponential term, A_{exp} , of the nucleation rate equation from experimental nucleation rate data at various

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temperatures. However, large discrepancies between the experimental pre-exponential term, A_{exp} and those predicted by the CNT, A_{theor} , were observed, e.g., [2,3]. However, it should be stressed that A_{exp} is very sensitive to the choice of the thermodynamic driving force for crystallization (the difference between the volume free energies of glass and crystal). Nevertheless, in every case, A_{exp} is higher than A_{theor} and the discrepancy may reach ~ 130 orders of magnitude.

Such a discrepancy can be avoided by calculating a specific surface energy from experimental nucleation rate data at each temperature, using the theoretical value of the pre-exponential term. Consequently, surface energy may be used as a fit parameter and its apparent temperature dependence can thus be obtained. This approach was first applied to the study of metallic systems [4,5] and then to silicate glasses [6,7]. Surface energy, calculated in such a way, slightly *increases* with temperature ($d\sigma/dT \sim (0.06\text{--}0.16) \times 10^{-3} \text{ J m}^{-2} \text{ K}^{-1}$). For instance, $d\sigma/dT = 0.059 \times 10^{-3}$ for Ga [4], 0.09×10^{-3} for Hg [5], 0.16×10^{-3} [6], 0.13×10^{-3} for L2S [7], 0.114×10^{-3} [3], 0.128×10^{-3} for N2C3S [8], $0.129 \times 10^{-3} \text{ J m}^{-2} \text{ K}^{-1}$ [this work] for 2NC3S; where L2S, N2C3S and 2NC3S denote glasses of stoichiometric composition $\text{Li}_2\text{O} \cdot 2\text{SiO}_2$, $\text{Na}_2\text{O} \cdot 2\text{CaO} \cdot 3\text{SiO}_2$ and $2\text{Na}_2\text{O} \cdot \text{CaO} \cdot 3\text{SiO}_2$.

Since all thermodynamic properties are temperature dependent, it is reasonable to assume that this also applies to surface energy. The temperature dependence of σ_∞ is also predicted by theory, see e.g., [9]. In particular, it has been shown [10] that $d\sigma_\infty/dT < 0$ when the molar volume of the liquid phase is higher than the molar volume of the crystal phase. A similar proposal was advanced in [11], i.e., crystal-melt surface energy should significantly increase with decreasing temperature, starting from the thermodynamic crystallization temperature. At temperatures below glass transition, σ_∞ could be almost as large as the surface energy of high-angle grain boundaries of the corresponding crystalline solid.

Moreover, there is experimental evidence that the liquid-gas surface tension of some glass-forming silicate melts decreases with increasing temperature [12]. According to [13], liquid-gas surface tension and crystal-liquid surface energy

are directly proportional to each other. Hence, one may expect that the latter also decreases with increasing temperature, contrary to the experimental results using the procedure described previously, i.e., fitting nucleation rate data to the CNT.

It should be stressed again that the surface energy $\sigma(T, R)$, calculated from nucleation data, refers to nuclei of critical size, R^* . The latter parameter, in turn, changes with temperature. Hence, as noted in [14], a $\sigma(T)$ plot may partially reflect a size effect. The issue of the size dependence of surface energy was elaborated in a number of papers, e.g., Refs. [14–17].

This paper focuses on the temperature dependence of the specific macroscopic surface energy, decoupling it from the size dependent part. Our aim is to calculate the surface energy of the crystal-melt interface as a function of temperature, $d\sigma_\infty/dT$, using crystal nucleation data for two silicate glasses and to compare it with theoretical predictions.

2. Theoretical background

According to the classical nucleation theory, e.g., [1], the homogeneous nucleation rate, I_{hom} , can be written as

$$I_{\text{hom}} = A_{\text{exp}} \left(-\frac{\Delta G_{\text{D}} + W^*}{kT} \right), \quad (1)$$

$$A = 2N_1 v_0 \left(\frac{\sigma a^2}{kT} \right)^{1/2},$$

where A is the pre-exponential term, k is Boltzmann's constant, $N_1 \cong 1/a^3$ is the number of structural (formula) units of melt with size a per unit volume, $v_0 = kT/h$ is the vibration frequency of a structural unit (for typical nucleation temperatures, $v_0 \cong 10^{13} \text{ s}^{-1}$), h is Planck's constant, σ is the nucleus/melt free energy per unit area, ΔG_{D} is the activation energy for transport of a structural unit across the nucleus/melt interface and W^* is the work of forming a critical nucleus.

From Eq. (1), the work of forming a critical nucleus (or the Gibbs free energy change due to the nucleus formation) determines, to a large

degree, the value of nucleation rate. It is well-known that if one neglects the strain energy associated to the formation of nuclei of size R (for simplicity, we assume nuclei of spherical shape) W can be written as

$$W = \frac{4}{3}\pi R^3 \Delta G_V - 4\pi R^2 \sigma, \quad (2)$$

where ΔG_V is the difference between the free energies of glass and crystal per unit volume of crystal. Fig. 1 shows the change of the Gibbs free energy as a function of the nucleus size for different approximations using typical values of the parameters given in the figure caption. The full curve (1) corresponds to the assumption of constant surface energy, σ and constant crystal density, ρ . However, as mentioned earlier, such assumptions may be incorrect in the case of nanometric size clusters. For example, regarding the description of a nucleus, Rusanov [9] claims that “A revision of the nucleation theories is required, considering the size dependence of the surface tension and of other thermodynamic parameters”. Thus, an account of the curvature dependence of the surface tension may result in significant

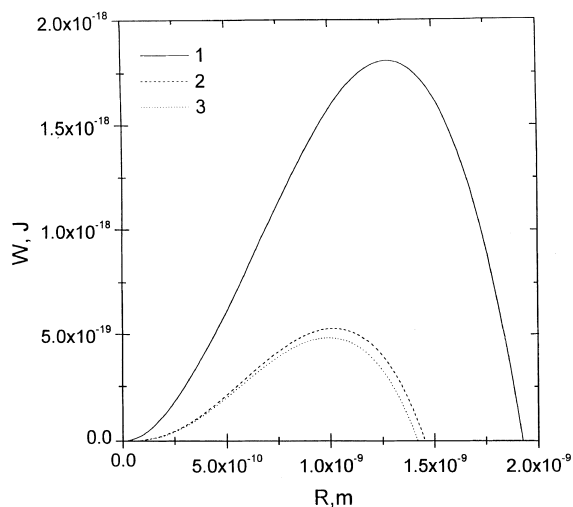


Fig. 1. Free energy change of the system due to the formation of a nucleus as a function of size. 1— σ and ρ are size independent; 2— σ is size dependent (Eq. (4)); 3— σ and ρ are size dependent. The following parameters were used in the calculations: $2\delta = 4.7 \times 10^{-10}$ m, $\Delta G_V = 4.08 \times 10^8$ J m $^{-3}$, $K = 5 \times 10^{-11}$ m 2 N $^{-1}$, $\sigma = 0.262$ J m $^{-2}$.

quantitative changes in the work of critical cluster formation.

There have even been suggestions that an account of this effect may result in a qualitative change of the shape of the thermodynamic barrier for nucleation, W^* [18]. A critical discussion of this point of view was given in [16].

Several approximate equations have been suggested for the curvature dependence of the surface energy, $\sigma = \sigma(R)$, e.g., [19–22]

$$\sigma(R) = \sigma_\infty \exp(-2\delta/R), \quad \text{Ref. [19]} \quad \text{Gibbs}, \quad (3)$$

$$\sigma(R) = \sigma_\infty / (1 + 2\delta/R), \quad \text{Ref. [20]} \quad \text{Tolman}, \quad (4)$$

$$\sigma(R) = \sigma_\infty (1 - \delta/R)^2, \quad \text{Ref. [21]} \quad \text{Rasmussen}, \quad (5)$$

$$\sigma(R) = \sigma_\infty (1 - 2\delta/R), \quad \text{Ref. [22]} \quad \text{Vogelsberger}, \quad (6)$$

where σ_∞ is the surface tension of a planar interface. Tolman's parameter, δ , characterizes the width of the interfacial region between the coexisting phases (it has the order of atomic dimensions). A general equation for curvature dependent surface tension was derived in [23], which includes Eqs. (3)–(6) as particular cases. Using an analysis of the limits of their validity, given in [16], we have chosen Tolman's Eq. (4) because it encompasses the smallest size (R) from which the equations are meaningful. Eq. (4) was derived for a liquid drop, assuming that $R \gg \delta$. It follows from our calculations (presented below) that the critical radius varies within a range of $1.7 < R^*/\delta < 4.6$. For this size range, Eq. (4) overestimates the surface tension by about 10–2.5%, respectively [20].

The change of the system's free energy due to nucleus formation, when a size dependence of surface energy is taken into account, is shown in Fig. 1 by a dashed curve (2). Here, Eq. (4) is used for a description of $\sigma(R)$.

Strictly speaking, the size dependence of the crystal density [$\rho(R) \cong \rho_\infty + \rho_\infty \kappa 2\sigma/R$, where κ is the isothermal compressibility and ρ_∞ is the crystal

density for $R \rightarrow \infty$] should also be taken into account [16]. For typical values of the thermodynamic parameters ($\kappa = 5 \times 10^{-6} \text{ atm}^{-1}$, critical radius $R^* = 10^{-9} \text{ m}$ and $\sigma_\infty = 0.2 \text{ J m}^{-2}$), the ratio between the thermodynamic driving force per unit volume of critical crystal and that of a macroscopic crystal, $\Delta G_V(R = R^*)/\Delta G_V(R \rightarrow \infty) = 1 + \kappa 2\sigma_\infty/R^*$, is only 1.013. On the other hand, if the crystal size varies from macroscopic to critical, the decreased surface tension leads to a decrease of the thermodynamic barrier of over three times (see Fig. 1, curves 1 and 2). The dotted curve (3) in Fig. 1 corresponds to the case where σ and ρ are size dependent. Thus, the use of a size dependent crystal density weakly affects the thermodynamic barrier of nucleation (compare curves 2 and 3). Hence, to a first approximation, we may treat the density of the nucleus as size independent.

One could also take into account the decrease of crystal density resulting from thermal expansion. Using the thermal expansion coefficient of polycrystalline lithium disilicate, $\alpha = 1.205 \times 10^{-5}$ (470–490°C) [24], the decrease in crystal density in the relevant temperature range for nucleation (T_g to $T_g + 100^\circ\text{C}$, where $T_g \sim 450^\circ\text{C}$ is the glass transition temperature) was estimated to be only 0.3%. Thus, we assume that the $\rho(T)$ dependence is too weak to affect the $\sigma(T)$ calculated from nucleation data.

Expressing the kinetic barrier of nucleation, ΔG_D , in terms of the induction period of nucleation, t_{ind} , or of the viscosity, η , Eq. (1) may be transformed into Eq. (7) or Eq. (8), respectively, see e.g., [8,25].

$$T_{\text{st}} = \frac{16}{3} \pi N_1^2 \frac{(kT)^{1/2} \sigma^{3/2}}{\Delta G_V^2 t_{\text{ind}}} \exp\left(-\frac{W^*}{kT}\right), \quad (7)$$

$$I_{\text{st}} = 2N_1 \left(\frac{\sigma a^2}{kT}\right)^{1/2} \frac{kT}{l^3 \eta} \exp\left(-\frac{W^*}{kT}\right), \quad (8)$$

where I_{st} is the steady-state homogeneous nucleation rate and l has the order of Si–O length.

The combination of Eqs. (2) and (4) results in

$$W = \frac{4}{3} \pi R^3 \Delta G_V - 4\pi \frac{R^3 \sigma_\infty}{R + 2\delta}. \quad (9)$$

From the condition $(\partial W / \partial R)_{R^*} = 0$ and taking the positive root one can find the critical radius R^*

$$R^* = \frac{(\sigma_\infty - 2\Delta G_V \delta) + (\sigma_\infty^2 + 2\Delta G_V \delta \sigma_\infty)^{1/2}}{\Delta G_V} \quad (10)$$

and then calculate the value of $W^* = W(R^*)$. Thus, on the right-hand side of Eqs. (9) and (10) and, correspondingly, of Eqs. (7) and (8), two parameters, crystal/liquid surface energy, σ_∞ and Tolman's parameter, δ , are unknown. The next part of the paper presents the results of fitting σ_∞ to experimental nucleation data for different (fixed) values of δ .

3. Calculations and results

Experimental nucleation rate data, induction times, viscosity and ΔG_V for $\text{Li}_2\text{O} \cdot 2\text{SiO}_2$ and $\text{Na}_2\text{O} \cdot 2\text{CaO} \cdot 3\text{SiO}_2$ glasses from [7,8,26,27] were used. These glasses exhibit internal, presumably homogeneous nucleation. Fig. 2 presents the temperature dependency of their steady-state

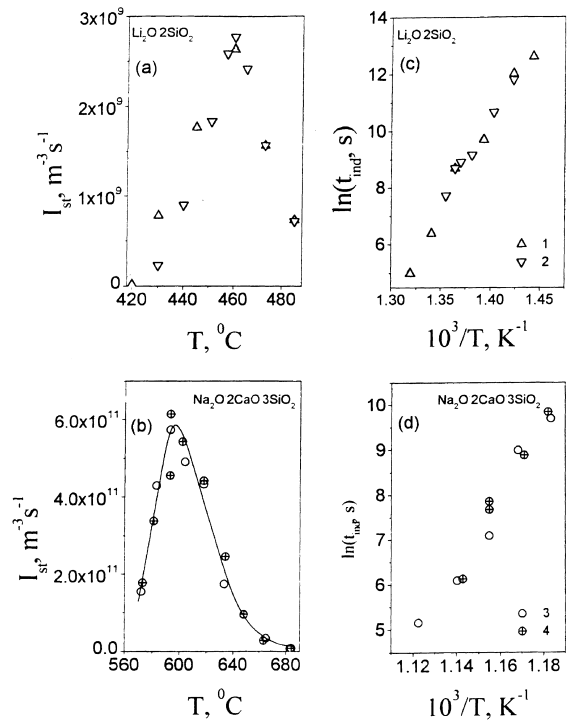


Fig. 2. Steady-state nucleation rates (a, b) and induction times for nucleation (c, d) as a function of temperature for L2S (a, c) and N2C3S (b, d) glasses. The data were taken from the following references: 1 – [26], 2 – [7], 3 – [8], 4 – [27].

nucleation rates (a, b) and induction times (c, d) vs. reciprocal temperature plots.

The values of I_{st} and t_{ind} , taken from the smoothed curves $I_{st}(T)$ (Fig. 2(a) and (b)) and from the Arrhenius plots $\ln(t_{ind}) \sim 1/T$ (Fig. 2(c) and (d)), were used for the calculations.

In the case of the N2C3S glass, ΔG_V was estimated by

$$\Delta G = -\frac{\Delta H_m}{T_m}(T_m - T) - \int_T^{T_m} \Delta C_p dT' + T \int_T^{T_m} \frac{\Delta C_p}{T'} dT', \quad (11)$$

where $\Delta G_V = -\Delta G/M$, ΔH_m is the latent heat of melting per mole and T_m is the melting point, $\Delta C_p = C_p^C - C_p^L (< 0)$ is the difference in specific heats between the crystalline and liquid phases and M is the molecular weight.

For the L2S glass, $\Delta G_V(\text{J m}^3) = 8.431 \times 10^8 - 548258.655T - 73.00247T^2 [T (\text{K})]$ was used. This expression fits the experimental data [28] well in the temperature range 693–758 K. Thermodynamic and viscosity data were taken from [29,30]. The following Fulcher equation was employed to describe the viscosity of L2S glass: $\log(\eta, \text{Pa s}) = -7.52 + 6259/(T - 406)$, $T (\text{K})$.

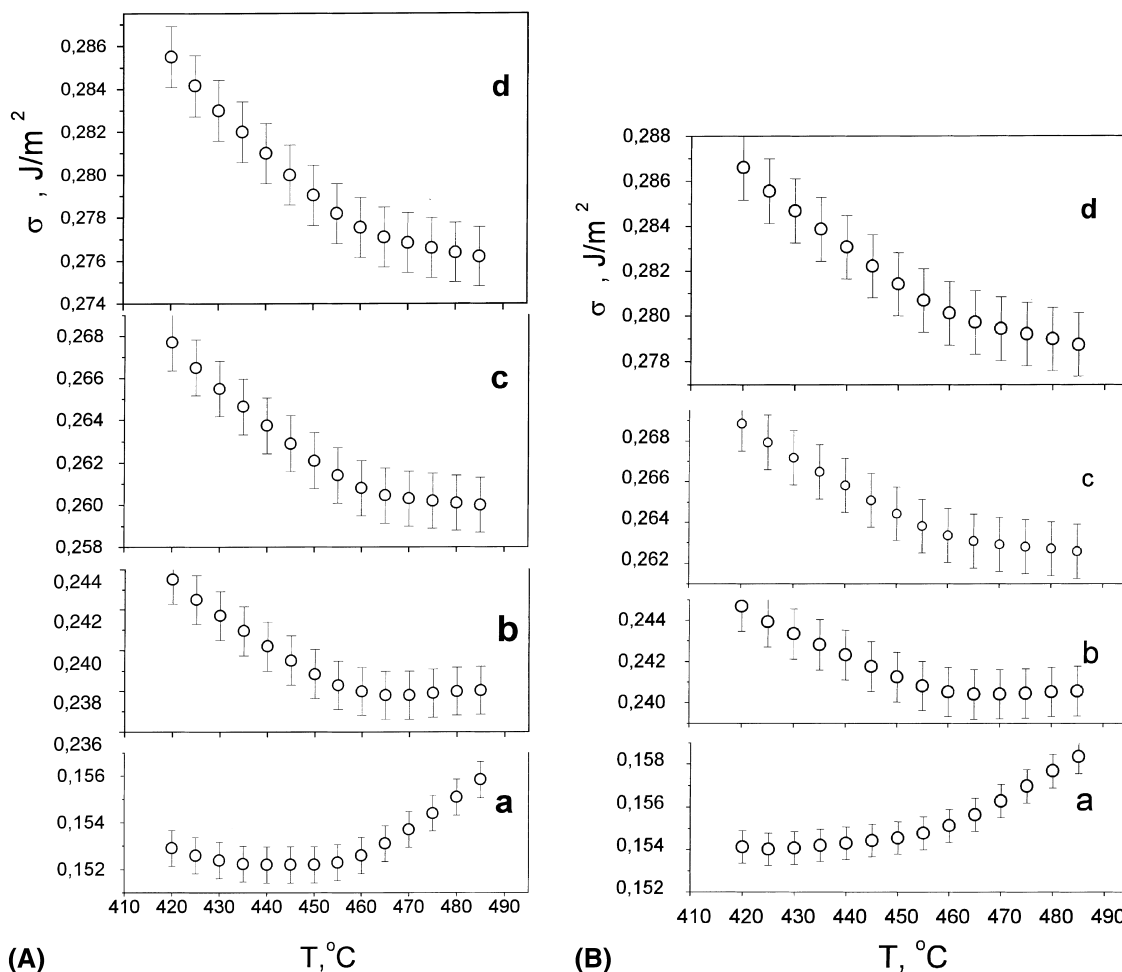


Fig. 3. (A) Crystal/ liquid surface energy for L2S glass calculated from Eq. (7) with different Tolman's parameters δ : 0 (a); 2.33×10^{-10} (b); 3×10^{-10} (c) and 3.5×10^{-10} m (d). (B) Crystal/ liquid surface energy for L2S glass calculated from Eq. (8) with different Tolman's parameters δ : 0 (a); 2.33×10^{-10} (b); 3×10^{-10} (c) and 3.5×10^{-10} m (d).

A commercial software was used for the numerical solution of Eqs. (4), (7) and (8) to obtain σ_∞ at different temperatures, using fixed values of Tolman's parameter, δ . Figs. 3 and 4 show the $\sigma_\infty(T)$ plots for L2S and N2C3S glasses, respectively. For those values of σ_∞ , the ratio between critical radius and Tolman's parameter ranges from $1.7 < R^*/\delta < 4.6$.

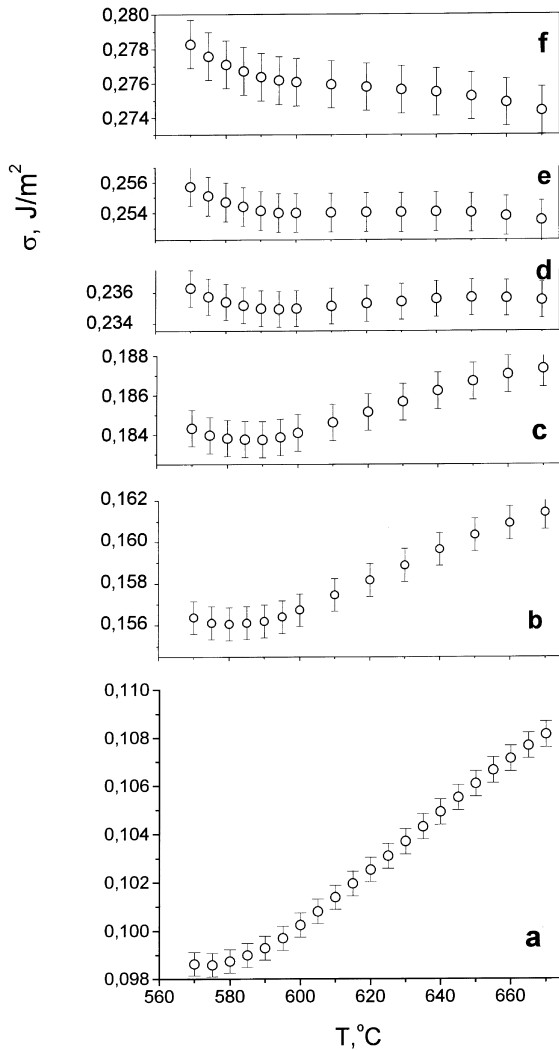


Fig. 4. Crystal/liquid surface energy for N2C3S glass calculated from Eq. (7), with different Tolman's parameters δ : 0 (a); 2.95×10^{-10} (b); 4.5×10^{-10} (c); 7.5 (d); 8.65×10^{-10} (e) and 10^{-9} m (f).

The numerical calculations allow one to find the values of surface energy for which the calculated nucleation rates equal the experimental ones (at a given set of Tolman's parameter and driving force). The experimental data were taken from the smoothed $I(T)$ curves. The errors in nucleation

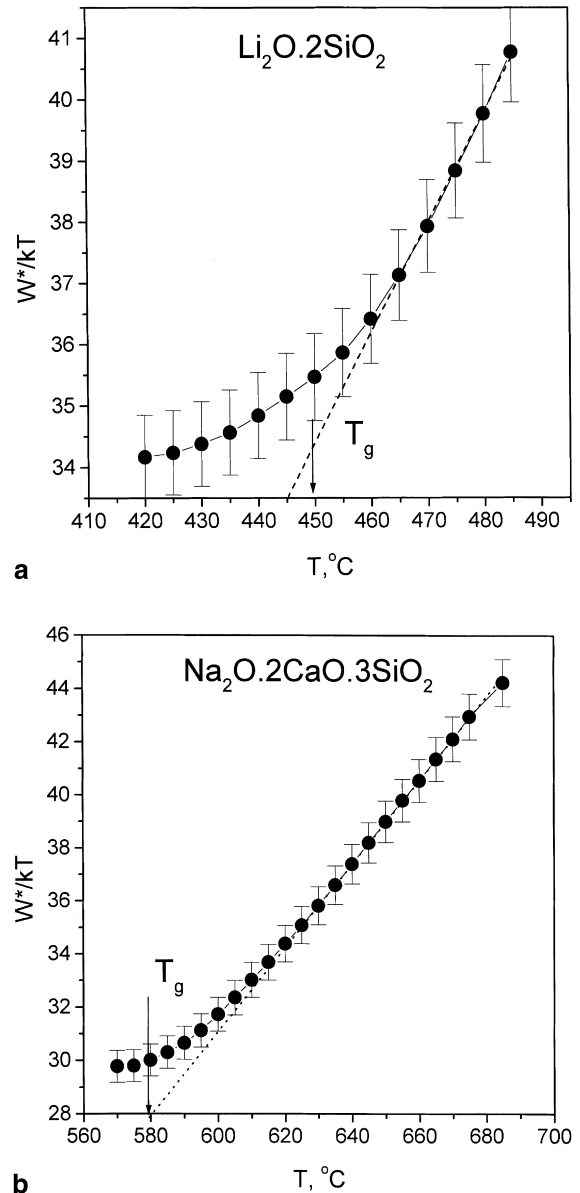


Fig. 5. (a) W^*/kT as a function of temperature for LS2 glass. (b) W^*/kT as a function of temperature for N2C3S glass.

rates were about 20%. In this case, W^*/kT varies $\sim 2\%$ and $\sigma \sim 0.5\%$ (shown in Figs. 3–5)

4. Discussion

As follows from Figs. 3 and 4, the calculation of $\sigma_\infty(T)$ is strongly affected by the choice of Tolman’s parameter. The case of $\delta = 0$ corresponds to a size independent surface energy, as employed by most authors. In this case, surface energy increases with temperature ($d\sigma_\infty/dT > 0$), from the glass transition temperature, T_g . At temperatures below T_g , surface energy changes very slightly (N2C3S glass) or is reduced to a minimum (L2S glass).

It is possible that, at these low temperatures, the elastic strains resulting from the difference between the densities of crystal and glass are important. In this case, the thermodynamic driving force for crystallization, ΔG_V , would be diminished by the elastic strain energy, which in turn would result in an increase of the thermodynamic barrier to nucleation, W^* . Fig. 5 gives circumstantial evidence for this assumption, i.e., the decrease of W^* slows down at temperatures close to T_g . Thus, when one uses overestimated values of ΔG_V (at low temperatures), one also obtains overestimated values of surface energy. This feature may explain the shape of $\sigma_\infty(T)$ dependence at low temperatures.

Fig. 6 shows the average values of $d\sigma_\infty/dT$ at $T \geq T_g$ vs. Tolman’s parameter. As δ increases, $d\sigma_\infty/dT$ reduces progressively and becomes negative at $\delta > 2.6 \times 10^{-10}$ m (L2S glass); 8×10^{-10} m (N2C3S glass). Thus, the Tolman parameter may be chosen so that the surface tension, σ_∞ , increases with decreasing temperature, in line with theory [10].

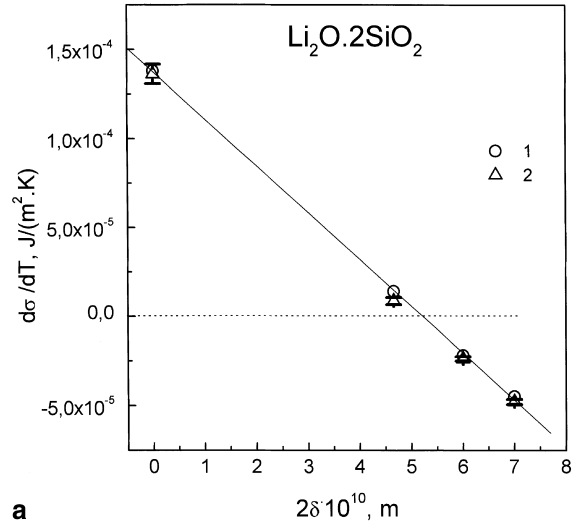
Moreover, one can estimate $(\sigma_m - \sigma(T_g))/(T_m - T_g) \cong d\sigma/dT$ using Eq. (22) from [10]

$$\sigma(T)/\sigma_m = 1 + (g/\gamma)(\varepsilon - 1)n_i\Delta G(T)/\Delta H_m, \quad (12)$$

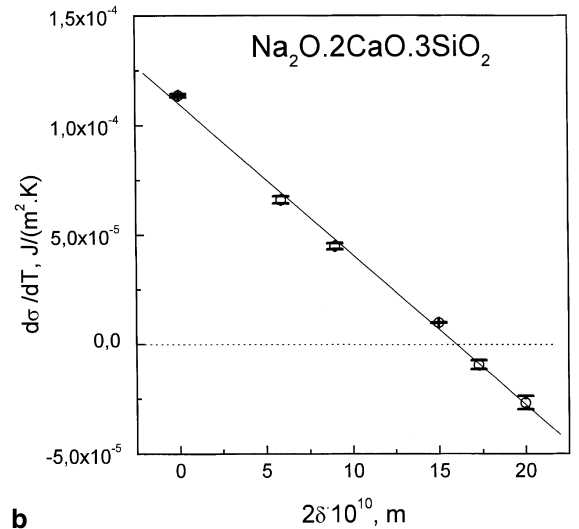
where γ is the coefficient in the well-known Scapski–Turnbull equation:

$$\sigma_m = \gamma\Delta H_m/v_c^{2/3}N_a^{1/3}, \quad (13)$$

where $\varepsilon = v_f/v_c(1 - g + gv_f/v_c)$, g is a numerical factor between 0 and 1, v_f and v_c the molar volumes of the melt and crystal, respectively and n_i is



a



b

Fig. 6. (a) $d\sigma_\infty/dT$ as a function of Tolman’s parameter for LS2. σ_∞ was calculated from Eq. (7) – 1 and Eq. (8) – 2. (b) $d\sigma_\infty/dT$ as a function of Tolman’s parameter for N2C3S. σ_∞ was calculated from Eq. (7).

the number of monolayers of the interface phase. For $v_f/v_c = 1.04$ (L2S), 1.02 (N2C3S), $\Delta G(T_g)/\Delta H_m = 0.42$ (L2S), 0.30 (N2C3S), $\gamma = 0.4$, $n_i = 2-4$ and $g = 0.5$, which provides the largest effect, $-d\sigma_\infty/dT$ varies from 2×10^{-5} to 4×10^{-5} $\text{J m}^{-2} \text{K}^{-1}$ for L2S glass and from 0.4×10^{-5} to 0.7×10^{-5} $\text{J m}^{-2} \text{K}^{-1}$ for N2C3S glass. These results are close to those calculated from the exper-

imental nucleation rates for $2.8 \times 10^{-10} < \delta < 3.4 \times 10^{-10}$ m (L2S) and $8.2 \times 10^{-10} < \delta < 8.5 \times 10^{-10}$ m (N2C3S) (Fig. 6).

Although these calculations are tentative, one can conclude that, by using reasonable values for Tolman's parameters, it is possible to obtain the temperature dependence of σ_∞ predicted by theory. Hence, the increase of the surface energy, $\sigma(R^*)$ with temperature, calculated by fitting nucleation rates to the CNT, is, indeed, caused by a size effect, namely, by the change of the critical size R^* . Thus, one may reach an agreement between an experiment and the CNT when the curvature dependence of the surface tension is taken into account.

5. Conclusions

The crystal/liquid surface energy σ_∞ was calculated from nucleation rate data for two silicate glasses, using Tolman's equation for size dependent surface tension. Reasonable values for the Tolman parameter may be chosen so that the surface tension decreases with temperature. The temperature dependence of surface energy is close to the theoretical predictions.

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